USING MOLES TO COUNT ATOMS

1. COUNTING THINGS
2. POPCORN – S, M, AND L
3. BUSHEL OF PEACHES
4. DOZEN OF EGGS
5. REAM OF PAPER
6. MASS MAY BE USED IN COUNTING
   * + 1. 10 GUMBALLS = 21.4 g
       2. 20 GUMBALLS = \_\_\_\_\_\_
       3. 30 GUMBALLS = \_\_\_\_\_\_
7. MOLES
8. THE SI BASE UNIT THAT DESCRIBES THE AMOUNT OF A SUBSTANCE.
9. USED TO COUNT A LARGE NUMBER OF SMALL PARTICLES.
10. AVOGADRO’S NUMBER = THE NUMBER OF PARTICLES/MOLE
11. = 6.22 X 1023 PARTICLES IN 1 MOLE OF A PURE SUBSTANCE
12. ONE MOL. OF POPCORN WOULD EQUIVALENT TO HIGH MANY KERNELS? \_\_\_\_\_\_\_\_
13. ONE MOL. OF GUMBALLS = \_\_\_\_\_\_\_\_\_\_\_ GUMBALLS.
14. WE CANNOT COUNT GUMBALLS AND POPCORN IN THIS MANNER.
15. SIMPLY TOO LARGE OF A UNIT TO COUNT POPCORN..
16. VERY USEFUL IN COUNTING ATOMS.
17. MOLES AND GRAMS ARE RELATED
18. MOLAR MASS – THE MASS IN GRAMS OF ONE MOLE OF A SUBSTANCE
19. ONE MOLE OF C-12 = 12.01 g
20. ONE MOLE OF SODIUM = \_\_\_\_\_\_\_ g
21. ONE MOLE OF CHLORINE = \_\_\_\_\_\_ g
22. ONE MOLE OF PHOSPHORUS = \_\_\_\_\_\_\_ g
23. MOLAR MASS IS THE SAME AS THE AVERAGE ATOMIC MASS OF A SUBSTANCE
24. ONE MOLE OF CARBON -12 CONTAINS 6.022 X 1023 ATOMS
25. 2 MOLES OF CARBON = \_\_\_\_\_\_g